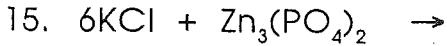
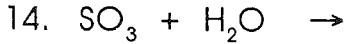
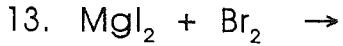
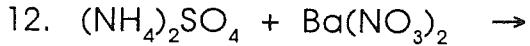
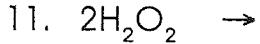
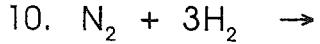
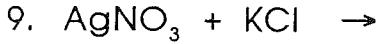
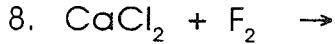
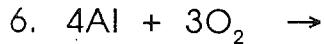
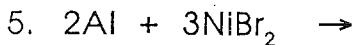
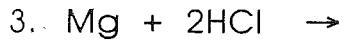


CLASSIFYING CHEMICAL REACTIONS

Name _____

Classify the following reactions as synthesis, decomposition, single replacement or double replacement.



Predicting Products:

Complete the equation given the reactants:

- a. $\text{Zn}(s) + \text{MgSO}_4(aq) \longrightarrow$
- b. $\text{Cd}(s) + \text{O}_2(g) \longrightarrow$
- c. $\text{HgO}(s) \longrightarrow$
- d. $\text{HCl}(aq) + \text{KOH}(aq) \longrightarrow$
- e. $\text{C}_5\text{H}_{12}(l) + \text{O}_2(g) \longrightarrow$
- f. $\text{Sr}(s) + \text{O}_2(g) \longrightarrow$
- g. $\text{Br}_2(aq) + \text{CaCl}_2(aq) \longrightarrow$
- h. $\text{Zn}(s) + \text{Ni}(\text{NO}_3)_2(aq) \longrightarrow$
- i. $\text{ZnSO}_4(aq) + \text{SrCl}_2(aq) \longrightarrow$
- j. $\text{AlCl}_3(aq) + \text{Na}_2\text{CO}_3(aq) \longrightarrow$
- k. $\text{Fe}(s) + \text{S}_8(s) \longrightarrow$
- l. $\text{C}_6\text{H}_6(l) + \text{O}_2(g) \longrightarrow$
- m. $\text{Pb}(s) + \text{KNO}_3(aq) \longrightarrow$
- n. $\text{HNO}_3(aq) + \text{Sr}(\text{OH})_2(aq) \longrightarrow$

Balance each equation in item 2.