Unit V Learning Log: Electrochemistry

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| **Learning Intentions** | **Practice** | **Evidence** | **Test Review** |
| G1: describe oxidation and reduction processes | V.1 and 2 p 189 – 195 #1 – 6POGIL Activity | * Balancing Complete Reactions Quiz Question #1
 | Multiple Choice: 1, 5, 7, 14, 15, 16, 17, 18, 25, 33, 37, 39, 41, 49, 50, 53, 61, 64, 65Written: 7, 21 |
| G2: analyse the relative strengths of reducing and oxidizing agents | V.3 p 190 – 200 #7 – 185 Stations activity  |  | Multiple Choice: 6, 19, 26, 27, 32, 34, 40, 42, 51, 52, 62, 63, 68Written: |
| G3: balance equations for redox reactions | V.4 and 5 p 201 – 207 #19 – 245 Stations activityWhiteboard group question | * Balancing Half-Reactions Quiz
* Balancing Complete Reactions Quiz Questions #2 – 4
 | Multiple Choice: 2, 3, 28, 29, 38, 54Written: 1, 2, 6, 8, 9, 10, 12, 18, 20, 22 |
| G4: determine the concentration of a species by performing a redox titration | V.7 p 210 – 214 #26 – 33Whiteboard group questions  | * Redox Titration Quiz
 | Multiple Choice: 67Written: 17 |
| H1: analyse an electrochemical cell in terms of its components and their functions | V.8, 9, and 10 p 215 – 228 #34 – 4721D DemonstrationGroup cell assignment  | * Electrochemical cell assignment
 | Multiple Choice: 4, 8, 12, 21, 22, 23, 30, 31, 36, 45, 46, 47, 48, 55, 56, 66Written: 3, 4, 5, 13, 16, 23 |
| H2: describe how electrochemical concepts can be used in various practical applications | V.11 p 228 – 233 #49 – 56 |  |  |
| H3: analyse the process of metal corrosion in electrochemical terms | V.12 p 233 – 236 #57 – 63 | * Corrosion Quiz
 | Multiple Choice: 43, 58, 69 |
| H4: analyse an electrolytic cell in terms of its components and their functions | V.13 p 237 – 243 #64 – 72 |  | Multiple Choice: 10, 11, 13, 20, 24, 44, 57, 59, 60, 71Written: 14, 15, 19 |
| H5: describe how electrolytic concepts can be used in various practical applications | V.13 p 243 – 246 #73 - 80 | * Electroplating Lab Activity
* Electrolysis Quiz
 | Multiple Choice: 9, 35, 70Written: 11 |