Using Ka and Kb in Calculations

**Remember ICE tables from Equilibrium**

1. An unknown weak acid, HW, is found to have a [H+] of 1.5 x 10-5 M when we mix a 2.5 M solution of the acid. Calculate Ka for this acid.
2. What is the [H+] in a 0.50 M CH3COOH solution?

\*you can assume x (in the denominator) is insignificant and ignore it in your calculation.

1. What is the [H3O+] in a 3.0 M solution of benzoic acid?

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